5-12-2014

CO₂ Reduction under Periodic Illumination of ZnS

Ruixin Zhou  
*University of Kentucky*, ruixin.zhou@uky.edu

Marcelo I. Guzman  
*University of Kentucky*, marcelo.guzman@uky.edu

Follow this and additional works at: [https://uknowledge.uky.edu/chemistry_facpub](https://uknowledge.uky.edu/chemistry_facpub)

Part of the Environmental Chemistry Commons, Materials Chemistry Commons, and the Physical Chemistry Commons

Right click to open a feedback form in a new tab to let us know how this document benefits you.

**Repository Citation**
Zhou, Ruixin and Guzman, Marcelo I., "CO₂ Reduction under Periodic Illumination of ZnS" (2014). *Chemistry Faculty Publications*. 23.  
[https://uknowledge.uky.edu/chemistry_facpub/23](https://uknowledge.uky.edu/chemistry_facpub/23)

This Article is brought to you for free and open access by the Chemistry at UKnowledge. It has been accepted for inclusion in Chemistry Faculty Publications by an authorized administrator of UKnowledge. For more information, please contact UKnowledge@lsv.uky.edu.
CO₂ Reduction under Periodic Illumination of ZnS

Digital Object Identifier (DOI)
http://dx.doi.org/10.1021/jp4126039

Notes/Citation Information

© 2014 American Chemical Society

This is an open access article published under an ACS AuthorChoice License, which permits copying and redistribution of the article or any adaptations for non-commercial purposes.
CO₂ Reduction under Periodic Illumination of ZnS
Ruixin Zhou and Marcelo I. Guzman*
Department of Chemistry, University of Kentucky, Lexington, Kentucky 40506, United States

ABSTRACT: The photoreduction of CO₂ to formate (HCOO⁻) in sphalerite (ZnS) aqueous suspensions is systematically studied in the presence of Na₂S hole scavenger. A series of cut-on filters at λcut-on ≥ 280, 295, 305, 320, and 400 nm are used to measure the reaction rate of formate production. The dependence of the measured reaction rates on λcut-on indicates that a wavelength of λ = 345 nm is associated with the actual bandgap of the semiconductor nanocrystals suspended in water. The results from apparent quantum yield measurements during periodic illumination experiments suggest that (1) valence-band holes on the surface of ZnS disappear within deciseconds due to the oxidation of the scavenger while simultaneously pumping electrons to the conduction band, (2) excited electrons in the conduction band of ZnS are transferred to CO₂ to produce the intermediate CO₂⁺, and (3) CO₂⁺ abstracts a proton from water and undergoes further photoreduction on the surface of ZnS in an overall time scale for steps 2 + 3 of a few milliseconds. The separation of both process merges at ∼29 ms because it decreases exponentially with a drop in [Na₂S] accompanied by a less negative surface potential. The behavior of the reaction rate at variable pH resembles the fraction of dissolved CO₂, discarding the direct participation of bicarbonate and carbonate in the reaction. Combined chromatographic, mass spectrometry, and spectroscopic studies provide new insights to understand the role of surface chemistry on the photoreduction of CO₂ on ZnS nanocrystals.

INTRODUCTION

Studying the reduction of CO₂ to organic compounds is an interesting problem that attracts the attention of groups that (1) aim to decrease the level of this abundant greenhouse gas in the atmosphere while simultaneously producing fuel feedstock, or (2) pursue a mechanism of prebiotic carbon fixation through abiotic reactions. The industrial scale conversion of CO₂ into fuels and the convenience of homogeneous or heterogeneous catalysis have been the subject of recent review articles. The photocatalytic reduction of CO₂ to HCOO⁻ through abiotic reactions has been the subject of recent review articles. The excellent reduction potential of ZnS and its low cost of production make it a viable candidate to promote CO₂ photoreduction in the photocatalyst for the reduction of CO₂ to HCOO⁻. However, limited progress has been made in this field due to the inertness of CO₂ and thermodynamic impediments. A typical reaction scheme for the photoreduction of CO₂ in ZnS colloidal suspensions involves (1) photoexcitation of ZnS to produce excited electrons (e⁻) and positive holes (h⁺) (reaction 1), which can be thermally loss (reaction 2), (2) photoexcited holes can be converted by H₂S hole scavenger to sulfate (reaction 3), and finally (3) adsorbed CO₂ molecules can accept an electron on the surface of ZnS to produce a radical anion intermediate (reaction 4) that is further reduced to HCOO⁻ (reactions 5):

$$\text{ZnS} \xrightarrow{hv} e^- + h^+$$  
$$e^- + h^+ \rightarrow \text{heat}$$  
$$\text{H}_2\text{S} + 4\text{H}_2\text{O} \rightarrow \text{SO}_4^{2-} + 10\text{H}^+ + 8e^-$$  
$$\text{CO}_2^{(aq)} + e^- \rightarrow \text{CO}_2^{*-}$$  
$$\text{CO}_2^{*-} + \text{H}^+ + e^- \rightarrow \text{HCOO}^-$$

The initial reduction of CO₂ proceeds through a single electron transfer to produce CO₂⁺, which is accompanied by a large energy redistribution from a linear molecule to a bent radical anion. The nucleophilic carbon in the surface-adsorbed radical anion abstracts a proton from water, forming a weaker acid–base pair than for the unviable oxygen radical. Finally, the HCOO radical accepts an electron to produce HCOO⁻. The overall two-electron transfer to convert CO₂ to HCOO⁻ requires a reduction potential $E^0 = -0.29$ V vs the normal hydrogen electrode (NHE) given in Table S1 (Supporting Information). In comparison, the conduction-band electron of ZnS is poised at $-1.04$ V vs NHE, and the
Experimental Details

Catalyst Preparation. Colloidal ZnS suspensions were prepared based on a previous method. The dropwise addition of 100 mL of 50 mM Na2S (99.1% assay, Sigma-Aldrich) to 100 mL of 50 mM ZnSO4 (ZnSO4·7H2O, EMD, 99.5% assay) took place under stirring and purging with N2(g) (UHP, Scott-Gross). Ultrapure water (18.2 MΩ cm, Elga Purelab Flex, Veolia) purged with N2(g) was used in every step. Sulfide concentrations were determined at 1:100 dilution in a 5% volume of 2 M NaOH (99.3% assay, Fisher Chemicals) using a sulfide ion selective electrode (Thermo, Orion 94-16). The potentiometric titration with a 1.000 M Cd(NO3)2 standard (cadmium nitrate tetrahydrate, Aldrich ≥99.0%) was used to standardize the sulfide stock employed in calibration curves. The ZnS colloidal suspension was centrifuged at 4400 rpm for 5 min (Eppendorf 5702) and rinsed twice with water to remove excess sulfide prior to irradiation.

Catalyst Characterization. Hydrodynamic particle size distributions and zeta potential (ζ) measurements were determined at variable pH using a DLS apparatus and ZetaPlus zeta potential analyzer (90 Plus/Bi-MAS, Brookhaven Instruments). For pH adjustment, 1 mL of fresh ZnS supernatant was diluted in 70 mL of 1 mM Na2S solution to bring the pH to 12, and then 0.02 M HCl (EMD Chemicals, GR ACS) was used to adjust the pH of ZnS suspensions. Undiluted samples were used for different measurements of ζ under variable [Na2S] at pH = 7.0. The surface potential of ZnS is taken as equivalent to ζ for unchanged permittivity and viscosity in the electrical double layer.

The absorption spectrum of scattering ZnS in water was obtained with a UV–vis spectrophotometer (Evolution 220, ISA-220 accessory, Thermo Scientific) using a built-in 10 mm silicon photodiode with a 60 nm Spectralon sphere. A 10 mm thick fused silica cuvette with water was used in the reference compartment in combination with a 1-mesh filter to subtract the background. The ISA-220 accessory was used in an alternative configuration to register the diffuse reflectance spectrum of dry ZnS powder as a Kubelka–Munk function against the certified Spectralon standard.

Photochemical Experiments. The colloidal suspensions for the photochemistry experiments were prepared just prior to use as indicated above. The addition of 0.5 mL of saturated Na2S stock solution provided the valence-band hole scavenger and prevented any adventitious O2 to enter the reactor. The adjusted Na2S concentration in the suspension was ca. 8 mM, and the loading of ZnS particles was 2.3 g L−1. The suspension pH was adjusted to 7 by bubbling CO2 (UHP, Scott-Gross) before sealing the vessel containing 200 mL of the suspension. The customized fused silica photochemical apparatus consisted of a 210 mL reaction vessel outfitted with a water jacket held at 288 K (Thermo Scientific bath circulator A25). For irradiation, a collimated 1 kW high-pressure Hg (Xe) arc lamp was used after removing infrared radiation with a water filter (Scheme S1, Supporting Information). Different optical filters (all from Newport) were used when indicated. The high loading of ZnS ensures the maximum saturation of light absorption. The 2.3 g L−1 of ZnS used in the photoreactor served as an efficient loading for the suspended particles to absorb more than 99% of the incident photon rate. The scattered photon rate measured with the detector at angles of 180° (forward scattered light), ±45°, ±90°, and ±135° is below 0.8% relative to the incident photon rate. Measurements of the photon flux at 330 nm, performed by ferrioxalate actinometry, were used to normalize the output of a calibrated air-cooled thermopile detector (Newport Corporation Model 818P-010-12) connected to an optical power meter (Newport Corporation Model 1918-C). This procedure has been previously demonstrated and made possible the use of the detector to measure the light intensity of individual experiments. Factory calibration of the thermopile performed at 1064 nm using NIST standards assures a linear response of the detector within a 4.0% error for the interval 190 ≤ λ ≤ 2100 nm, which is automatically adjusted by a correction factor preloaded in the meter program. The values of ΦHCOO− at λ = 330 nm were calculated from measurements of RHCOO− and I0 using a series of filters: (1) a neutral density filter transmitting only 34% of the light, (2) a broad band optical filter absorbing in the infrared, (3) a cut-on filter at 320 nm, and (4) a band-path filter transmitting only λ = 330 ± 10 nm. Because the measurements do not account for photon scattering, Φ values are lower limits of the quantum yields. The generation of light pulses was possible by means of a combined mechanical shutter and optical chopper (Newport 75159) set by a controller (Newport model 75160) with different blades (models 75162, 75163, and 75164). The shutter mode was used for pulsing between 0.02 and 0.5 Hz (1−25 s on period) and the shutter mode for the 10 and 2400 Hz regime (0.208−50 ms on period). A second chopper and controller (SRS model SR540) was used for pulsing in the intermediate range, between 1 and 10 Hz (0.05−0.5 s on period).

Analysis of Products. After irradiation for a selected time period, 5 mL of suspension was withdrawn from the reaction vessel through a septum, centrifuged (5 min at 4400 rpm), and the supernatant was filtered (IC Acrodisc 0.2 μm pore size; Pall Corp.). Analyses of products in the filtrate and in samples processed by solid-phase extraction (Hypersep C18, 1 mL, 100 mg) were performed by ion chromatography (IC). The second water eluent fraction of SPE cartridges preconditioned with
water and methanol were injected with an autosampler (Dionex AS) into the IC system (Dionex ICS-2000). The IC was equipped with an IonPac AS11-HC analytical column (2 × 250 mm), an IonPac AG11-HC guard column (2 × 50 mm), an ASRS-300 (2 mm) suppression module, an eluent generator (EGC III) with KOH cartridge, an anion trap column (CR-ATC), and a conductivity detector used to match the retention times of standards. The gradient separation applied a flow of 0.38 mL min⁻¹ with an initial mobile phase of 1 mM KOH for 8 min followed by three linear increases of (1) 1.4 mM min⁻¹ up to 15 mM, (2) 1.5 mM min⁻¹ up to 30 mM, and (3) 3 mM min⁻¹ up to 60 mM. After IC separation, a mass spectrometer detector (Thermo MSQ Plus) interfaced with an electrospray ionization (ESI) probe (450 °C and 70 psi N₂, nebulizing gas, 1.9 kV needle voltage, and 50 V cone voltage, operated in the negative mode) identified the mass-to-charge ratio (m/z) of species. An auxiliary pump (Lab Alliance) provided a flow of 0.12 mL min⁻¹ methanol (Fisher Optima) mixed through a T connector with the eluent flow before aerosolization in the ESI chamber. Chromeleon and Excalibur software were used to control and process data from both systems. Concentrations were obtained from calibration curves and/or using the method of standard addition. Standards included solutions of sodium formate (Acros Organics, 99%), sodium sulfate (Fisher Chemicals, 99.9% assay) in 1% methanol, sodium sulfate (Fisher Chemicals, 99.9% assay), and sodium thiosulfate (EM Science, 97% assay).

Sulfide concentrations were potentiometrically quantified with the calibrated sulfide ion selective electrode described above. The decay of sulfide was monitored during periodic illumination experiments at λ = 330 nm to extract the decay rate constant (k, s⁻¹). Samples withdrawn from the reactor were immediately centrifuged at 4400 rpm for 5 min, and the supernatant was treated with an equal volume of 2.0 M NaOH. The time series of sulfide decay followed first-order kinetics with correlation coefficients r² ≥ 0.990.

## RESULTS AND DISCUSSION

### ZnS Characterization

The characterization of the aqueous suspensions of ZnS by dynamic light scattering (DLS), measurements of zeta potential (ζ), and the UV–vis absorption spectrum included below provides direct information to understand the surface photoreactions of interest. The findings in this paper indicate that surface species dominate the reduction of CO₂ and the oxidation of the hole scavenger.

Noninvasive DLS reports particle sizes such as those present during the experiments in aqueous conditions (Figure 1), avoiding filtering and drying preparation steps needed in XRD and TEM, which could result in the oxidation and change in the surface charge of ZnS. The effect of acidification on particle size and surface charge in ZnS colloidal suspensions—sequential addition of acid reduced the pH of these suspensions from 12.0 to 2.3—is also shown in Figure 1. The measurements applied DLS to compare the diffusion of particles in the colloidal suspension. In addition, the crystallite size D from the XRD pattern (Figure S1, Supporting Information) and transmission electron microscopy (TEM) images are not exactly the same of the particle size obtained by DLS (Figure 1) because ZnS particles in suspension are made up of several different crystallites.

The intensity correlation function of each sample was measured over three runs at a scattering angle of 90° from the incident laser (λ = 658 nm) beam. Particles with size between 35 and 950 nm were detected at pH 10 and 4, respectively (Figure 1A). The particle size is described by a unimodal distribution for the whole pH range. The synthesized ZnS at 20 °C initially experiences slight changes of aggregate size, but the distribution remains unimodal with diameters of 40–200 nm down to pH 5.9. It should be noted that further acidification led to considerable larger particles for the pH range 3.1–5.1 and peaks at pH = 4.0 to a value of 940 nm. No considerable difference in the particle size distribution was observed for ZnS synthesized at 5, 20, or 40 °C.

The negative values for the zeta potential in Figure 1B with magnitude larger than the boundary of ~30 mV, in the range of pH 7.0–12.0, indicate that the particles of ZnS repeal each other, and there is no tendency to flocculate. As the pH drops below 5.9, the suspension becomes less stable and the zeta potential becomes increasingly less than ~30 mV; thus, there is no force to prevent the particles coming together and flocculation occurs. The continuous addition of dilute HCl eventually neutralizes the negative charge at the isoelectric point of pH = 4.2 ± 0.1 obtained by interpolation (red dashed line in Figure 1B). Further addition of acid results in a buildup of positive charge. For the syntheses at 5 and 40 °C, the isoelectric point is within the ±0.1 range of pH measured at 20 °C.

The characterization of dry ZnS by other techniques could be affected by processes occurring in bulk and aggregated catalyst. This information for X-ray diffraction (XRD) patterns (Figure S1) and Raman spectroscopy (Figure S2) during the reaction, transmission electron microscopy (TEM) and energy-dispersive X-ray spectroscopy (EDS) (Figure S3), and Brunauer–Emmett–Teller (BET) measurements can be found in the Supporting Information.

### Identification of Products

Figure 2 shows the species identified as anions in the chromatogram for irradiation at λ > 200 nm. Peaks elute in the following order: formate (HCOO⁻, m/z = 45), bicarbonate (HCO₃⁻, m/z = 61), bisulfite (HSO₃⁻, m/z = 81), bisulfate (HSO₄⁻, m/z = 97), hydrogen thiosulfate (H₂S₂O₇⁻, m/z = 113), and hydrogen dithionate (H₂S₄O₆⁻, m/z = 161). The same products were observed for irradiation with the different cut-off filters (λcut-off ≥ 280, 295, 305, and 320 nm) and for monochromatic light (λ = 330 nm). Control experiments that demonstrate the heterogeneous photo-
production of formate on the surface of the ZnS semiconductor were carried out in the presence of one or more of colloid (ZnS), ultraviolet light (hν; λ > 200 nm), carbon dioxide (CO₂), and sulfur-based hole scavenger (S₈). In the absence of ZnS, there is no production of formate (control A). The negative result for the production of formate under dark conditions (control B) indicates that a thermal mechanism is not the cause of this reaction. No formate production was observed unless CO₂ as an electron acceptor (control C) and sulfide hole scavenger as an electron donor (control D) were present.

The oxidation of sulfide hole scavenger during the reaction also produces a small amount of yellow colloid visually observed during irradiation and identified as elemental (rhombic) sulfur, S₈, in the Raman spectrum (Figure S2, Supporting Information). This species is not derived from the catalyst that remains intact as shown by XRD and TEM data (Supporting Information). Rhombic sulfur is an oxidation intermediate produced by the sulfur-based hole scavenger. The simultaneous photooxidation and substitution mechanism below protects the integrity of ZnS, consumes solution-phase sulfide ions in the system, and produces elemental sulfur.²⁶

\[
\text{ZnS} + 2h^+ \rightarrow \text{Zn}^{2+} + S \tag{5}
\]

\[
\text{Zn}^{2+} + \text{HS}^- \rightarrow \text{ZnS} + \text{S} + \text{H}^+ \tag{6}
\]

\[
2h^+ + \text{HS}^- \rightarrow \text{S} + \text{H}^+ \tag{7}
\]

Given the dissociation of H₂S (pKₐ = 7.05 and pKₐ₂ = 14.92 at 25 °C),²⁷ both available acidic species at pH = 7 (dissociation fraction αₐ₁,H₂S = 0.53 and αₐ₂,H₂S = 0.47) may participate as initial hole scavenger. However, neutral H₂S should be the main reductant because HS⁻ is repelled by the negatively charged surface of ZnS for higher pH than the isoelectric point of ZnS (pH = 4.2, Figure 1). CO₂ is reduced to HCOO⁻ in the conduction band of ZnS, while all involved sulfur species participate in oxidation reactions at the valence band. S₈, S₂O₃²⁻, SO₃²⁻, and SO₄²⁻ (Scheme 1) are formed as intermediate products during the oxidation of sulfide toward SO₄²⁻ with the possible transfer of electrons summarized in Table S2 (Supporting Information).

Upon dissolution of gaseous CO₂ in water at 15 °C (Henry’s law constant Kₕ = [CO₂(aq)]/P_CO₂ = 0.0260 mol L⁻¹ atm⁻¹ at 15 °C)²⁸ several species coexist in equilibrium:

\[
\text{CO}_2(g) \rightleftharpoons \text{CO}_2(aq) \tag{8}
\]

\[
\text{CO}_2(aq) + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{CO}_3 \tag{9}
\]
H$_2$CO$_3$ → HCO$_3^-$ + H$^+$, \( pK_a = 6.42^{30} \)  

HCO$_3^-$ → CO$_3^{2-}$ + H$^+$, \( pK_a = 10.43^{30} \)  

and only a negligible amount of all dissolved CO$_2$ (CO$_2$(aq)) + H$_2$CO$_3$) is in the form H$_2$CO$_3$. The species present in the previous equilibrium system (Figure 4) depend on pH as follows: (1) At pH < \( pK_a \) = 6.42 dissolved CO$_2$ is dominant and (2) at pH > \( pK_a \) = 10.43 carbonate is dominant, while (3) for intermediate pH values (\( pK_a \) < pH < \( pK_a \)) bicarbonate is the main species. The dependence of \( R_{\text{HCOO}^-} \) on pH (Figure 4) under full lamp irradiation resembles the fraction of total dissolved CO$_2$ present in equilibrium, indicating that CO$_2$ is the only species that controls the process. The charged species (HCO$_3^-$ and CO$_3^{2-}$) are not primary reagents undergoing photoreduction.

This observation agrees with a model of reactive surface species adsorbed to the negatively charged surface of ZnS (Langmuir–Hinshelwood mechanism), which prefers the interaction with neutral dissolved CO$_2$ over the electrostatic repulsion to HCO$_3^-$ and CO$_3^{2-}$ anions. However, an Eley–Rideal mechanism cannot be completely discarded.$^{31–33}$

A narrow particle size distribution is confirmed by the sharp drop observed in the reflectance spectrum of dry ZnS powder with reflection edge at \( \lambda = 300 \) nm (Figure 5). The associated valence band of the dry ZnS powder has a bandgap of 3.76 eV (\( \lambda = 330 \) nm), which is blue-shifted from the absorption edge of wet (bulk) ZnS at \( \lambda = 345 \) nm (Figure 5). The absorption edge is centered at \( \lambda = 312 \) nm in the ZnS suspension in water (Figure 5). Although the particles could experience quantum size effects,$^{34}$ the experiments take place in water, and the absorption spectrum of ZnS colloidal suspensions is needed to estimate the bandgap energy under actual experimental conditions. The slopes in Figure 3 give the \( R_{\text{HCOO}^-} \) values plotted in Figure 5 vs specific \( \lambda_{\text{cut-on}} \) of irradiation. The bandgap of ZnS colloidal suspensions can be better determined from the extrapolated straight line in Figure 5.

This is the first time that such a determination of the bandgap is made based on the measurement of reaction rates and irradiation wavelengths. The equation for the linear regression in Figure 5 is \( R_{\text{HCOO}^-} (\mu \text{M min}^{-1}) = 14.152 - 0.0410 \times \lambda_{\text{cut-on}} \) (nm) with a coefficient of correlation \( r^2 = 0.995. \) The reaction rate falls to zero \( (R_{\text{HCOO}^-} = 0) \) for an extrapolated cut-on wavelength \( \lambda_{\text{cut-on}} = 345 \) nm. The value of \( \lambda_{\text{cut-on}} = 345 \) nm is in excellent agreement with the absorption spectrum for the ZnS colloidal suspension (red line in Figure 5), registered from the supernatant remaining in solution after centrifugation of the 2.3 g L$^{-1}$ ZnS colloidal suspension (DLS yields \( D = 51 \pm 2 \) nm). Remarkably, the absorption spectrum in Figure 5 falls to zero above the same \( \lambda_{\text{cut-on}} \). The bandgap energy for ZnS suspension in water is calculated using the equation \( E_{\text{BG}} = h\lambda_{\text{cut-on}} \) = 5.74 $\times$ 10$^{-19}$ J $\equiv$ 3.59 eV, where \( h = 6.626 \times 10^{-34} \) J s is the Plank constant and \( \lambda = 3.00 \) $\times$ 10$^{-8}$ m is the speed of light. The measured \( E_{\text{BG}} \) for nanoparticles in water is within the range of previous reports for ZnS,$^{13}$ and it does not display quantum size effects, as expected for diameters larger than the Bohr radius of ZnS (2.50 nm).$^{35}$

For continuous monochromatic irradiation at \( \lambda = 330 \) nm, the ratio of \( R_{\text{HCOO}^-} = 5.92 \times 10^{-9} \) mol L$^{-1}$ s$^{-1}$ to \( I_0 = 4.10 \times 10^{-7} \) einstein L$^{-1}$ s$^{-1}$ (measured by actinometry), yields \( \Phi_{\text{HCOO}^-} = 1.44\% \) as a lower limit of the quantum yield, because light is extinguished both by absorption of the photocatalyst and by light scattering from the particles in suspension.$^{24}$ This monochromatic \( \Phi_{\text{HCOO}^-} \) should not be compared to those larger values also reported by ferrioxalate actinometry (e.g., for ZnS and MnS$^{15,36}$ employing broadband irradiation (200–400 nm). Care must also be taken when comparing our measurement to the higher \( \Phi_{\text{HCOO}^-} = 20\% \) at \( \lambda = 280 \) nm.$^{37}$ More importantly, several factors could affect \( \Phi_{\text{HCOO}^-} \) in different experimental setups, making the comparison of \( \Phi_{\text{HCOO}^-} \) by different groups difficult.$^{38}$ Particularly, experimental systems could have different (1) intensity of monochromatic illumination, (2) pH, (3) mass transport (e.g., stirring rate in the reactor), and (4) cell geometry.$^{39}$ In addition, the structure and associated catalyst properties could vary to considerably affect the observed \( \Phi_{\text{HCOO}^-}. \) For example, whether the experimental preparation skipped the final steps after the synthesis (centrifugation, washing, and rinsing of the catalyst with water), the reaction rate at \( \lambda = 330 \) nm increased to \( R_{\text{HCOO}^-} = 3.72 \times 10^{-8} \) M s$^{-1}$, representing a 6.3 times more efficient process with \( \Phi_{\text{HCOO}^-} = 9.07\% \). The \( \Phi_{\text{HCOO}^-} \) values reported are useful to compare the set of reactions occurring under similar conditions for the specific structure characterized. Overall, the information derived from the \( \Phi_{\text{HCOO}^-} \) values in the section below is of general application to the photocatalyst and enabled the determination of the lifetime of surface redox carriers.

**Lifetime of Redox Carriers on the Surface of Photoexcited ZnS.** Figure 6A shows the changes in \( \Phi_{\text{HCOO}^-} \) (at pH 7.0) under periodic illumination with alternate and equally lasting bright (\( \tau_L \)) and dark (\( \tau_D \)) intervals. In Figure 6A, for the upper limit chopper frequencies examined, \( \Phi_{\tau_L} \) merges with...
the value measured under continuous illumination ($\Phi_{\text{HCOO}^-} = 1.44\%$). However, two well-resolved inflection points are observed when transitioning from $\Phi_{\text{L} \rightarrow \text{ox}}$ to the lower limit chopper frequencies ($\Phi_{\text{L} \rightarrow \text{red}}$), which correspond to redox carriers with different reactivity.39 The two inflection points fitted with a nonlinear least-squares regression for a double sigmoid curve occur at $t_1 = 3.3$ ms and $t_2 = 0.13$ s (Figure 6A) for $[\text{Na}_2\text{S}]_0 = 2.0$ mM and pH = 7.0 and can be associated with the transfer of reducing conduct-band electrons and the oxidizing valence holes acting over sulfide, respectively. Close values of $t_1 = 2.4$ ms and $t_2 = 0.14$ s (to those obtained by measuring $\Phi_{\text{HCOO}^-} = R_{\text{HCOO}^-}/I_0$) are potentiometrically determined from the first-order rate constant for sulfide decay ($k_{\text{H}_2\text{S}}$), a stable species in dark controls (Figure 6B).

Experiments at variable initial $[\text{Na}_2\text{S}]$ are useful to investigate which transition corresponds to oxidizing and reducing species. The measured surface area of ZnS assures experimental conditions of submonolayer coverage by the hole scavenger. The previous imbalance of oxidizing valence holes acting over sulfide creates an excess redox potential ($E^\circ$) and surface charge. The previous imbalance affects the half-reduction and half-oxidation rate constants of ZnS $k_{\text{red}} = k_{\text{e}}e^{-[\alpha_{\text{red}}F/RT]}(E-E^\circ)$ and $k_{\text{ox}} = k_{\text{ox}}e^{[\alpha_{\text{red}}F/RT]}(E-E^\circ)$, respectively, where $E^\circ$ is the standard redox potential, $R$ is the gas constant, $F$ is the Faraday constant, $T$ is the absolute temperature, $n_{\text{red}}$ and $n_{\text{ox}}$ are the number of electrons transferred in each half-reaction, $k_{0}$ is the standard heterogeneous rate constant, and $\alpha_{\text{red}}$ and $\alpha_{\text{ox}}$ are the dimensionless charge transfer coefficients.39 In our experiments, the potential is affected by changing the concentration of the reducing species in the range $0.50 \leq [\text{Na}_2\text{S}]_0 \leq 16.9$ mM, what modifies the surface potential of ZnS (Figure S5, Supporting Information) as represented by $\zeta$ for unchanged permittivity and viscosity in the electrical double layer.8,10 Figure 7A shows, as expected, both transitions $t_1$ and $t_2$ have opposite trends with increasing $[\text{Na}_2\text{S}]$. Analogously, Figure 7B indicates the same kind of dependence for the transitions on $\zeta$. The experimental information in Figure 7A demonstrates how the two photoredox transitions observed are related to the rate constants $k_{\text{red}}$ and $k_{\text{ox}}$, which are reciprocal to $t_1$ and $t_2$, respectively: $t_1 \propto k_{\text{red}}^{-1}$ and $t_2 \propto k_{\text{ox}}^{-1}$. Given the previous inverse relationships, and the positive exponential form of $k_{\text{ox}}$, an increase in the concentration of hole scavenger (or the potential) predicts an increase in the value of $t_2$ with $[\text{Na}_2\text{S}]$ in Figures 7A and its decrease in Figure 7B with $\zeta$. The intersection point of both regression lines in Figure 7B predicts that both transitions become indistinguishable for a surface potential of $-16$ mV at 29 ms. On the contrary, the measured $t_1$ value decreases with $[\text{Na}_2\text{S}]$ and increases for larger $\zeta$ (Figure 7B) due to the negative exponential form describing $k_{\text{red}}$. This interpretation of $t_1$ being related to $k_{\text{red}}$ is in agreement with the recently reported higher rate of reduction for $\text{H}_2(\text{g})$ generation in CdS nanocrystals.41 The higher rate of $\text{H}_2(\text{g})$ production observed followed an exponential function for more negative potentials.41

## CONCLUSIONS

Remarkably, the reactive intermediates in the photoexcited semiconductor exist for relative long times,39 as observed here for ZnS$^*$. Characteristic surface carriers remain active for less than 10 and 215 ms for the transfer of mobile electrons and loss of fixed holes, respectively. In other words, the reactive carriers can be detected for longer time scales than the range explored (e.g., 150 fs $\leq \tau_{1/2} < 100$ $\mu$s)32 during time-resolved spectroscopy studies of semiconductors. The photoreduction of $\text{CO}_2$ to $\text{HCOO}^-$ in aqueous colloidal suspensions of ZnS in the presence of $\text{Na}_2\text{S}$ directly consumes dissolved $\text{CO}_2$ as the reactive species adsorbed to the surface of ZnS. A new method to report the bandgap of the semiconductor in the actual
hydrated state is demonstrated based on the dependence of the reaction rate on the wavelength of irradiation. The measurement of quantum yields under periodic illumination offers mechanistic information that provides the time scale \((10^{-3} - 10^{-1}\) s) of redox processes on the surface of ZnS, which occur at several orders of magnitude slower than excitation of ZnS. Further studies of this system are underway with the goals of comparing catalyst stability, improving efficiency, and replacing the sacrificial electron donor by other species.

**REFERENCES**


